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Question 2 $2 \text{ Fe(s)} + 3 \text{ O}_2\text{(g)} \rightarrow \text{Fe}_2\text{O}_3\text{(s)}$ $\Delta H^\circ = -824 \text{ kJ mol}^{-1}$

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 $\text{Ag}_2\text{CrO}_4(\text{s}) \rightleftharpoons 2 \text{Ag}^+(\text{aq}) + \text{CrO}_4^{2-}(\text{aq})$ $K_{\text{sp}} = 2.6 \times 10^{-12}$ at 25°C

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$\text{CH}_3\text{CH}_2\text{COOH}(\text{aq}) + \text{H}_2\text{O}(\text{l}) \rightleftharpoons \text{CH}_3\text{CH}_2\text{COO}^-(\text{aq}) + \text{H}_3\text{O}^+(\text{aq})$ Propanoic acid, $\text{CH}_3\text{CH}_2\text{COOH}$, is a carboxylic acid that reacts with water according to the equation above. At 25°C the pH of a 50.0 mL sample of 0.20 M $\text{CH}_3\text{CH}_2\text{COOH}$

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