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FULL CHAPTER || CLASS 10

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Solution: The reaction of the acid + base gives a product of salt + water, which is considered as neutralization reaction.

Examples: $\text{NaOH} + \text{HCl} \rightarrow \text{NaCl} + \text{H}_2\text{O}$

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chapter 2 Acids, bases and
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Short Answer Type. Question
1. Two solutions A and B
have pH values of 5 and 8
respectively. Which solution
will be basic in nature ...

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Ans:- Salts are classified
into four categories on the
basis of the strength of
acid and base. they are:
Salt of strong acid and
strong base.e.g NaCl, KCl;
Salt of strong acid and weak
base. e.g. NH_4Cl , NH_4I ;
Salt of weak acid and strong
base. e.g. CH_3COONa , COO^-
 Na^+ + Salt of weak acid and
weak base. e.g. CH_3COO^-
 NH_4^+ + How is common salt
found in nature?

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2. Bases react with acids to produce salts and water; water to produce acids and salts; salts to produce acids and water; neither acids, salts, nor water

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The salts of strong acids and weak bases give acidic solution having pH less than 7. Example, NH_4Cl , Ammonium Chloride will have pH less than 7. The salts of weak acids and strong bases give basic solution having pH more than 7. Example, Na_2CO_3 , Sodium Carbonate will have pH more than 7.
Question 7.

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Answer A reaction in which an acid and base react with each other to give a salt and water is termed as neutralization reaction. For Example: (i) $\text{NaOH} + \text{HCl} \rightarrow \text{NaCl} + \text{H}_2\text{O}$ (ii) $\text{HNO}_3 + \text{KOH} \rightarrow \text{KNO}_3 + \text{H}_2\text{O}$ 15. Give two important uses of washing soda and baking soda. Answer Two important uses of washing soda are:

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Answer. (a) HCl is acid and NaOH is base whose combination forms the common salt. Its formula is NaCl (Sodium chloride). It is

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Chemistry
obtained from sea water. (b)
Rock salt is the common name
for the mineral "halite".
Its chemical formula is
NaCl. It may be white or
light blue or yellow
depending upon impurities
present in it.

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Chapter Wise Important
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(A) inorganic acid (B)
organic acid (C) soft
organic acid (D) strong
inorganic acid. Answer.

Answer: (a) 5. When milk of
magnesia reacts with acetic
acid it produces (A) basic
salts (B) acidic salts (C)
neutral salt (D) complex
salt. Answer. Answer: (c) 6.
Which of the following
phenomena will occur when a
small amount of acid is
added to water ...

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Turmeric is a natural indicator. On adding its paste to acid and base separately, which colours would be observed. (a) Yellow in both acid and base (b) Yellow in acid and red in base (c) Pink in acid and

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~~Chemistry~~
yellow in base (d) Red in
acid and blue in base.

Answer. Answer: (b) Yellow
in acid and red in base

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Question 1.

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Acids, Bases and Salts In this first quiz on GCSE Chemistry we take a look at one vital part of the subject - acids, alkalis, bases and salts. In it we find out the difference between alkalis and bases, how acids and alkalis react with one another, how neutralisation can produce salts and the relevance of pH values.

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| Revise pH Values and
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Acids, bases and salts An acid is a compound that is

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Chemistry defined by its physical and chemical properties. Acids taste sour and react with metals and polyatomic ions called carbonates. A carbonate is a charged cluster of Carbon and Oxygen atoms.

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Acid + Base \rightarrow Salt +
Water. Question 3: Rain
becomes acidic due to.
carbon dioxide. sulphur
dioxide. nitrogen dioxide.
all of the above. Correct
Option is : 4. Solution :
Rain becomes acidic because
of carbon dioxide, sulphur
dioxide and nitrogen dioxide.
They dissolve in rain drops
to form carbonic acid,
sulphuric acid and nitric
acid ...

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