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CHAPTER 16 REVIEW Reaction Energy SECTION 1 SHORT ANSWER Answer the following questions in the space provided. 1. For elements in their standard state, the value of ΔH_f° is . 2. The formation and decomposition of water can be represented by the following thermochemical equations: $\text{H}_2(\text{g}) + \frac{1}{2}\text{O}_2(\text{g}) \rightarrow \text{H}_2\text{O}(\text{g})$ 241.8 kJ/mol $\text{H}_2\text{O}(\text{l})$ 241.8 kJ/mol ? $\text{H}_2(\text{g}) + \frac{1}{2}\text{O}_2(\text{g})$

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CHAPTER 16 REVIEW Reaction Energy SECTION 1 SHORT ANSWER Answer the following questions in the space

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provided. 1. For elements in their standard state, the value of ΔH_f° is 0. 2. The formation and decomposition of water can be represented by the following thermochemical equations: $\text{H}_2(\text{g}) + \frac{1}{2}\text{O}_2(\text{g}) \rightarrow \text{H}_2\text{O}(\text{g})$ $\Delta H = -241.8 \text{ kJ/mol}$ $\text{H}_2\text{O}(\text{l}) \rightarrow \text{H}_2(\text{g}) + \frac{1}{2}\text{O}_2(\text{g})$ $\Delta H = +241.8 \text{ kJ/mol}$...

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Title: Chapter 16 Review Reaction Energy Section 1 Short Answer Author: Nicole Fassbinder Subject: Chapter 16 Review Reaction Energy Section 1 Short Answer

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the study of the transfers of energy as heat that accompany chemical reactions and physical changes.

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Section 16.1 Collision Theory: A Model for the Reaction Process. Goals. To describe a model, called collision theory, that helps us to visualize the process of many chemical reactions. To use collision theory to explain why not all collisions between possible reactants lead to products. To use collision theory to explain why possible reactants must collide with an energy equal to or above a certain amount to have the possibility of reacting and forming products.

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Chapter 16 - The Process of Chemical Reactions

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