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Pure Substances and Mixtures, Elements \u0026amp;

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3.2.1/3.2.2 Describe the differences between elements, compounds and mixtures. ~~States of Matter : Solid Liquid Gas 10 Amazing Experiments with Water~~ Changes in the State of Matter - Science Experiments - Solid to Liquid, Liquid to Gas, Gas to Liquid Changes of States of Matter Conservation of

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Mass in Chemical Reactions The States of Matter - Changes of State - Solid Liquid Gas Plasma | Physical Science | Gen Chemistry States of Matter *Understanding Atoms, elements, and molecules Part #1 (9min)* Study Jams - Properties of Matter *States of Matter and Changes of State - Science for Kids*

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A change in physical form of a substance or physical properties of a substance like color, melting point, state, size, shape, and arrangement of molecules is known as physical change. A change in chemical properties of a substance like composition, reactivity with water, corrosiveness, and pH is known as chemical change. In chemical change, a new substance will be formed.

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145 CHAPTER 9 SOLUTIONS MANUAL 35. Aqueous solutions of potassium iodide and silver nitrate are mixed, forming the precipitate silver iodide. chemical equation:  $KI(aq) + AgNO_3(aq) \rightarrow KNO_3(aq) + AgI(s)$  complete ionic equation:  $K^+(aq) + I^-(aq) + Ag^+(aq) + NO_3^-(aq) \rightarrow K^+(aq) + NO_3^-(aq) + AgI(s)$

## **Chemical Reactions**

Solution. Physical: boiling and melting are physical changes.

When water boils, no bonds are broken or formed. The

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change could be written:  $\text{H}_2\text{O} \left( l \right) \rightarrow \text{H}_2\text{O} \left( g \right)$  Chemical: The dark grey nail changes color to form an orange flaky substance (the rust); this must be a chemical change.

## 3.6: Changes in Matter: Physical and Chemical Changes

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240 Chemistry: Matter and Change • Chapter 12 Solutions Manual CHAPTER 12 SOLUTIONS MANUAL 30. Compare and contrast sublimation and evaporation. In both processes, the substances become a vapor. During sublimation, the substance goes from the solid phase directly to the vapor phase. During evaporation, particles in a liquid gain enough

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## **States of Matter**

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