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Reduction-oxidation reactions are known as redox reactions.

An oxidising agent causes a substance to be oxidised, that is to lose electrons, it itself gains these electrons and so becomes reduced e.g. hydrogen peroxide, iodine solution. A reducing agent

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Oxidation and
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Oxidation in terms of
electron transfer.
Oxidation is the loss
of electrons; The
substance which

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loses electrons is said to be oxidised;

Oxidising agent: This is the substance

which causes another substance to be

oxidised i.e. it takes electrons; Examples

of oxidising agents.

Hydrogen Peroxide (H_2O_2) □ Used to bleach hair

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□ Oxidation & Reduction □ oxidation and reduction, oxidation numbers, rules for oxidation numbers, balancing redox equations

About Tara: Tara has been teaching Higher Level Chemistry at The Institute of Education since 2001. During this time many

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of her students have
achieved excellent
results with one of her
Reactions

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The whole subject can seem a bit far-fetched and hard to relate to. Honestly it gets easier the more you learn, especially topics like organic chemistry and stoichiometry and so on. The basics you

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learn in the first few chapters are really relevant down the line, particularly in chapters like redox titrations and so on.

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educators throughout
the world are
attempting to make
students aware of
how mankind benefits
from the advances
being made

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Redox Reactions
Redox Titrations The
most common
oxidising agent used
for volumetric analysis
is potassium
permanganate. It
reacts with, and is
used for the
determination of
reducing agents, the
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example being iron(II)

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published by the
Department mean
that for 2021,
students will get away
with knowing a slightly
lesser portion of the
course. □ Students
must answer on any
eight questions out of
eleven, as before, but
it is no longer
mandatory to answer
two out of three
questions in Section A

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The oxidation numbers tell us how electrons are divided up or shared between atoms in a chemical compound. The oxidation numbers also tell us how electrons move in an

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There are a set a rules that we use to determine oxidation number. Group 1A elements (alkali metals) always have an oxidation of +1.

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