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1. Aqueous solutions of acids have a sour taste.
2. Acids change the color of acid-base indicators.
3. Some acids react with active metals and release hydrogen gas, H_2
4. Acids react with bases to produce salts and water.
5. Acids conduct electric ...

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Modern Chemistry Chapter 14 Acids & Bases
Section 1 Properties of Acids & Bases Aqueous
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3 is the Lewis acid because it accepts an electron pair in forming a covalent bond. It is not a Brønsted-Lowry acid, because it is not donating a proton. e. Which of the reactants is the Lewis base? Explain your answer. NF_3 is the Lewis base because it donates an electron pair in forming a covalent bond. 120 ACIDS AND BASES MODERN CHEMISTRY

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CHAPTER 14 REVIEW . Acids and Bases. SHORT ANSWER Answer the following questions in the space provided. 1. a. Write the two equations that show the two-stage ionization of sulfurous acid in water. stage 1: $\text{H}_2\text{SO}_3(\text{aq}) + \text{H}_2\text{O}(\text{l}) \rightleftharpoons \text{H}_3\text{O}^+(\text{aq}) + \text{HSO}_3^-(\text{aq})$ b. Which stage of ionization usually produces more ions? Explain your answer.

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binary acid. An acid that contains only 2 different elements: hydrogen and one more electronegative elements. oxyacid. An acid that is a compound of hydrogen, oxygen, and a third element, usually a nonmetal. Arrhenius acid. A chemical compound that increases the

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concentration of hydrogen ions, H^+ , in aqueous solution.

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an acid that is a compound of hydrogen, oxygen, and a third element, usually a nonmetal. arrhenius acid. a substance that increases the concentration of hydronium ions in aqueous solution. strong acid. an acid that ionizes completely in a solvent. weak acid.

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