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#### CHAPTER 9 REVIEW Stoichiometry SECTION 3 PROBLEMS

Write the answer on the line to the left. Show all your work in the space provided.

- 88% The actual yield of a reaction is 22 g and the theoretical yield is 25 g. Calculate the percentage yield.
- 6.0 mol of  $N_2$  are mixed with 12.0 mol of  $H_2$  according to the following equation:  
$$N_2(g) + 3H_2(g) \rightarrow 2NH_3(g)$$

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branch of chemistry that deals with mass relationships in  
compounds and in chemical reactions.  $2 \text{ HgO} \rightarrow 2 \text{ Hg} + \text{O}_2$   
2- a) 2 mol HgO 2 mol Hg & 2 mol HgO 1 mol O<sub>2</sub>...

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6 Chemical Bonding

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