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The relative masses were obtained by multiplying the atomic ratios and atomic masses. You can see that a sample of N_2O weighing 44.02 grams contains 28.02 g of nitrogen and 16.00 g of oxygen. The mass percent of each element is calculated from its relative mass divided by the sum of the relative masses. Chemical compounds with integral atomic ratios, like nitrous oxide, are described as stoichiometric compounds, and they permit many simple calculations.

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Study Guide – Unit 8: Stoichiometry Vocabulary – Give definitions and examples where applicable
1. Product
2. Reactant / Reagent
3. Theoretical Yield
4. Actual Yield
5.

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Percent Yield 6. Excess Reagent 7. Limiting Reagent 8. Law of Conservation of Matter Short Answer – explain the following: 9. What is conserved in every balanced equation (multiple answers)? 10.

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Stoichiometry Study Guide Prerequisite: Chemistry 1102

Credit Value: 1 Text: Chemistry. Mustoe, Jansen, et al;

McGraw-Hill Ryerson; 2004. Chemistry Concentration

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The best definition for stoichiometry is the simple one: it's a way to figure out how much stuff you're going to make in a chemical reaction, or how much stuff you'll need to make a chemical reaction do what you want. When we put it that way, stoichiometry isn't so bad. We can deal with the crazy name if it's that simple.

Stoichiometry Introduction | Shmoop
From aluminum to xenon, we explain the properties and composition of the substances that make up all matter.

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In your textbook, read about stoichiometry and the balanced equation. For each statement below, write true or false. _____ true _____ 1. The study of the quantitative relationships between the amounts of reactants used and the amounts of products formed by a chemical reaction is called stoichiometry.

_____ true _____ 2.

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Stoichiometry Questions. 1. Copper has two stable isotopes $^{63}_{29}\text{Cu}$ (atomic mass = 62.93 amu; 69.09% natural abundance) and $^{65}_{29}\text{Cu}$ (atomic mass = 64.9278 amu; 30.91 % natural abundance). Calculate the average atomic mass of copper. 2. How many atoms of sulfur are in 25.1 g of S? 3. Calculate the molecular mass (in amu) of caffeine ($\text{C}_8\text{H}_{10}\text{N}_4\text{O}_2$). 4.

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study guide 4.0 Unit 4 Overview: Chemical Reactions

According to the College Board, "This unit explores chemical

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transformations of matter by building on the physical transformations studied in Unit 3. Chemical changes involve the making and breaking of chemical bonds.

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